

Unit 2, Video 2: Isotopes

1. True or False: All atoms of an element must have the same number of neutrons.
2. True or False: Carbon-12 has a total of 12 protons plus neutrons.
3. True or False: The mass number can be determined if the atomic number and number of neutrons are provided.
4. True or False: The mass number can be determined by looking at the periodic table.
5. Boron has two common isotopes: boron-10 and boron-11. Which do you expect to be the most abundant?
6. Fill in the missing information from the table below:

Isotope Symbol		Bromine-80
Atomic Number	20	
Mass Number		
Protons		
Neutrons	22	

7. Use the information below to calculate the average atomic mass of the unidentified element.

	Mass	Percent Abundance
Isotope 1	34.96885	75.78%
Isotope 2	36.96590	24.22%

8. Based on your calculations, what element is described in question 7?