

## Unit 2, Video 8: Orbital Notation

1. How many electrons are in the 2s sublevel for oxygen in its ground state?



2. The orbital notation for oxygen is depicted above. Describe how the Pauli Exclusion Principle was obeyed in this example.
3. Using the orbital notation of oxygen shown above, describe how the Aufbau Principle was obeyed.
4. Using the orbital notation of oxygen shown above, describe how the Hund's Rule was obeyed.
5. How many orbitals should be drawn for a P sublevel?
6. In the space below, draw the orbital notation for the elements carbon and chlorine.