

Unit 3, Video 3: Periodic Trends In Ions

1. Increasing the number of charged particles within an atom means increasing the number of _____ and _____.
2. True or False: The first electron removed from an atom should come from the outside energy level.
3. According to the ionization energy trend, which of the following elements would you expect to have the highest ionization energy? Sulfur, argon, oxygen, or neon?
4. True or False: The concept of shielding makes an electron on an inner energy level easier to remove than an electron in the outermost energy level.
5. According to the trend on the periodic table, which of the following elements has the highest electronegativity value? Sodium, chlorine, potassium, bromine.
6. True or False: The atom has an outer barrier making the size easy to measure.
7. According to the trend in atomic radius on the periodic table, which of the following elements would you expect to have the smallest radius? Magnesium, sulfur, calcium, selenium?
8. True or False: Fluorine is more reactive than bromine.
9. Which would you expect to be larger? Calcium or the Calcium Ion?