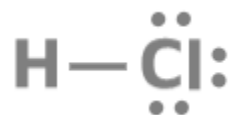
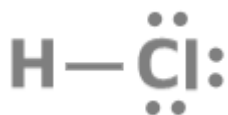


Unit 4, Video 8: Polar Covalent Bonds

1. Describe the trend for electronegativity found on the periodic table.
2. True or False: Polar covalent bonds are made up of atoms with differences in their electronegativity values.
3. The bond between hydrogen and chlorine is a polar covalent bond. Use the Lewis Structures below to demonstrate the two ways this polarity can be indicated.



4. Describe what is happening with the electrons in the bond between hydrogen and chlorine.
5. True or False: All molecules that contain a polar covalent bond are polar molecules.
6. List the molecular shapes that often cancel out polarity in molecules.
7. For each of the molecules below, indicate if the molecule is polar or nonpolar.

