

Unit 5, Video 14: Limiting Reactants

1. What is a limiting reactant?
2. What is an excess reactant?
3. Describe how a question can be recognized as a limiting reactant problem.
4. The steps necessary for solving the question below are listed. Use the space provided to answer the question. Be sure to show your work.



Given the chemical reaction above, how many grams of ammonia can be produced when 10.0 grams of nitrogen and 10.0 grams of hydrogen react?

Step 1: Convert to moles.

Step 2: Divide by the coefficients.

Step 3: Use the limiting reactant to answer the original question.